For th	e following species:
$BeF_2$	
BCl <sub>3</sub>	
CCl <sub>4</sub>	
PBr <sub>5</sub>	
$SI_6$	
$\mathrm{BH_2}^-$	
$NI_3$	
ClF <sub>4</sub> <sup>+</sup>	
$\mathrm{SF}_5^-$	
-	Draw the best Lewis structure
	Pradict the electron domain geometry

- Predict the electron domain geometry
- Predict the molecular geometry
- Predict the polarity of the molecule

Species	Lewis Structure	Electron Domain Geometry	Molecular Geometry	Polarity
BeF <sub>2</sub>	<b>:F.</b> ──Be ── <b>F:</b>	linear	linear	Nonpolar
BCl <sub>3</sub>	: ; ; ; ; ; ; ; ; ; ; ; ; ; ; ; ; ; ; ;	trigonal planar	trigonal planar	Nonpolar
CCl <sub>4</sub>	::: ::: ::: ::::	tetrahedral	tetrahedral	Nonpolar
PBr <sub>5</sub>		trigonal bipyramidal	trigonal bipyramidal	Nonpolar
SI <sub>6</sub>		octahedral	octahedral	Nonpolar
BH <sub>2</sub> <sup>-</sup>	<mark>н — в — н</mark>	trigonal planar	bent	Polar

NI <sub>3</sub>	II:	tetrahedral	trigonal pyramidal	Polar
ClF <sub>4</sub> <sup>+</sup>	F. C1 (1+)	trigonal bipyramidal	see saw	Polar
SF <sub>5</sub>	:F: F: F:	octahedral	square pyramidal	Polar